

Legendre Polynomials

Bernd Schröder

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 arises when the equation $\Delta u = f(\rho)u$ is solved with separation of variables in spherical coordinates. (QM: hydrogen atom!) The function $y(\cos(\phi))$ describes the polar part of the solution of $\Delta u = f(\rho)u$.

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3. The solutions of both equations *must* be finite on $[-1, 1]$.
4. Because 0 is an ordinary point of the equation, it is natural to attempt a series solution.

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$$c_{k+2} = \frac{k(k-1) + 2k - \lambda}{(k+2)(k+1)}c_k = \frac{k(k+1) - \lambda}{(k+2)(k+1)}c_k.$$

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9. So the solutions we are interested in will be polynomials with even powers (for $\lambda = l(l+1)$ and l even) or polynomials with odd powers (for $\lambda = l(l+1)$ and l odd).

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Choosing $c_0 = (-1)^{\frac{l}{2}} \frac{l!}{2^l \left[\left(\frac{l}{2}\right)! \right]^2}$ gives

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which is the customary way the Legendre polynomials are stated.

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(The generalized Legendre equation is good reading.)

Electron Orbitals

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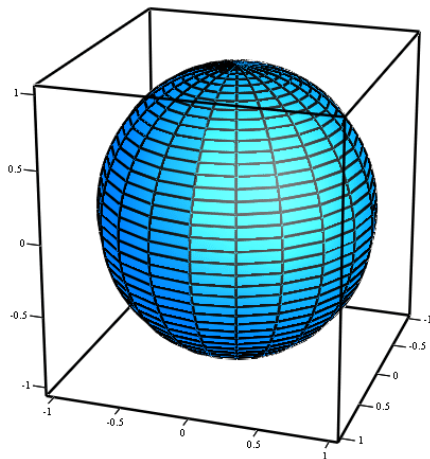
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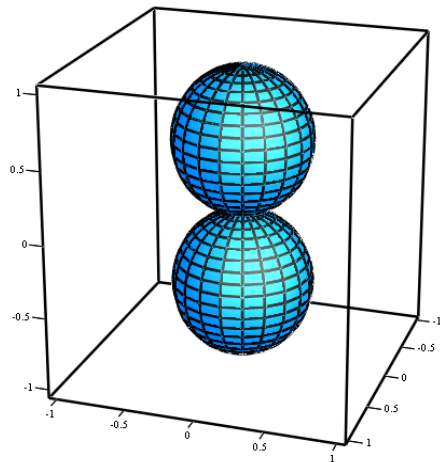
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6. “Shape” must be carefully interpreted. Large values for $\rho(\phi) = P_n(\phi)$ in the picture indicate a large probability (density) that the electron’s location’s polar angle is around the angle ϕ .

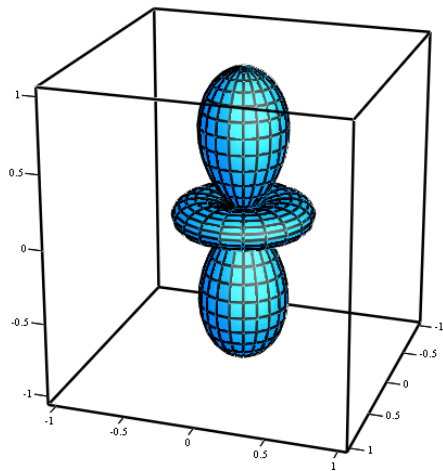
$$\rho = |P_0(\cos(\phi))|: 1s \text{ Orbital}$$



$\rho = |P_1(\cos(\phi))|$: 2p Orbital



$\rho = |P_2(\cos(\phi))|$: 3d Orbital



$$\rho = |P_3(\cos(\phi))|: 4f \text{ Orbital}$$

